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Basic

Introduction,

Mole to Mole,

Grams to Grams,

Mole Ratio

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9: Stoichiometry
examples

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130 Chapter 9 WP

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Stoichiometry
Example 4 Heat
of Reaction Ch 9
Section 9.2:

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~~Magic Number~~
~~Method~~

The more general

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~~Stoichiometry~~
principle,
beyond quantum

9.2 Ideal

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~~explained +~~

~~Crash Chemistry
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~~Problem: Mass~~
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~~Limiting~~

~~Reactant and~~

~~Excess Reactant~~

~~9.1 Introduction~~

~~to Stoichiometry~~

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~~Limiting \u0026amp; Excess Reactant, Theoretical \u0026amp; Percent Yield~~

~~Chemistry Mole Ratio Practice Problems~~

**CHEMISTRY -- CH.
9 TEST REVIEW**

Balancing
Chemical
Equations
Practice

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Problems Chapter

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Answer CHAPTER 9
REVIEW
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SECTION 2

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

1. 4.5 mol The following equation represents a laboratory preparation for oxygen gas:

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$2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$
How many moles
of O_2 form if
3.0 mol of

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Answers Reaction

stoichiometry

uses molar

relationships to

determine the

amounts of

unknown

reactants or

products from

the amounts of

known reactants

or products.

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fewer steps are
required to
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problems when.
... Chemistry

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Answer Key

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9-1 Introduction
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Composition

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deals with mass

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relationships of
elements in
compounds
Reaction

Stoichiometry -
Involves mass
relationships
between
reactants and
products in a
chemical
reaction I.
Reaction
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Problems A. Four
problem Types,
One Common
Solution

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Stoichiometry
CHAPTER 9 REVIEW
Stoichiometry
MIXED REVIEW
SHORT ANSWER
Answer the
following
questions in the

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Stoichiometry.
space provided.

1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$
- a. What is the value of the coefficient x in this equation? 4
- b. What is the molar mass of C_3H_4 ? 40.07 g/mol
- 2 mol

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O 2:1 mol H 2O
c. What is the
mole ratio of O
2 to H

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Review Answers

reaction. Define

reactant. Define

product.

Identify the

products and

reactants in a

reaction.

Identify a

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Stoichiometry.
Relate the
symbols in a
chemical
equation to the
words in a word
equation. Write
the word
equation from a
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Review #1 – #18,
#31, & #38

Answers . 38. To ensure that all magnesium is converted to MgO , I would use pure oxygen, not air, to carry

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Stoichiometry

out the reaction,
because Mg could
react with N_2 in
air to form

Mg_3N_2 . The pure
oxygen should be
in excess. 5. a.

50 mol HI 6. a.
15.8 Holt

Chemistry

Chapter 9:

Stoichiometry -

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Answers

Chapter 9

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Multiple Choice

Answers

Chapter 9:

Standard Review

Worksheet 1.

Answers will

vary. An example

is included

below: $2\text{H}_2\text{O}_2$

(aq) $2\text{H}_2\text{O}(l) +$

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0 2 (g) This describes the decomposition reaction of hydrogen peroxide.

Microscopic: Two molecules of hydrogen peroxide (in aqueous solution) decompose to produce two

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molecules of
liquid water and
one molecule of
oxygen gas.

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Standard Review
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SHORT ANSWER

Answer the following questions in the space provided.

1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ 4 a. What is the value of the coefficient x in this

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equation? 40.07
g/mol b. What is
the molar

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