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Basic*

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Equilibrium:

Crash Course

Chemistry #28

~~Ice Table~~

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~~Constant~~

~~Expression,~~

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~~K_p , K_c ,~~

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~~Examples How To~~

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To Calculate The

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Pressures -

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Chemical
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*Dissolution Of
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(Part 3) ~~How to
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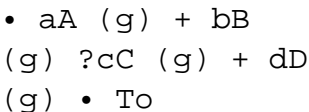
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Equilibrium

Outline 1.

N_2O_4 - NO_2

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Determination of

K 4.

Applications of

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Effect of
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$2\text{HI}(\text{g}) \rightleftharpoons \text{H}_2(\text{g}) + \text{I}_2(\text{g})$ If pure $\text{HI}(\text{g})$ is placed in a sealed container at 52°C , $\text{H}_2(\text{g})$ and $\text{I}_2(\text{g})$ are formed. Then the reverse reaction can occur forming $\text{HI}(\text{g})$. The partial pressure of

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HI(g) drops, slowing the forward reaction, all the while, the partial pressures of H₂(g) and I₂(g) increase; increasing the rate ...

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concentrations
of the species
remain constant
over time. if
the initial
concentrations
of the reactants
and products
change they will
return to the

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concentrations
at equilibrium.

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equilibrium

constant will

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equilibrium

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equation. $2A + 3B$

$\Rightarrow C$. $K = (C_p) /$

$(A_p)^2 * (B_p)^3$.

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$(P_{\text{partial}})(V) = (n)(0.08206)(T)$

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Equilibrium are
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7.1 EQUILIBRIUM IN PHYSICAL PROCESSES

The characteristics of system at equilibrium are better understood if we examine some physical processes. The most familiar examples are

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phase

transformation
processes, e.g.,
solid ? liquid

liquid ? gas

solid ? gas

7.1.1 Solid-

Liquid

Equilibrium Ice

and water kept

in a perfectly

insulated

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