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Using your  
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theory of acids  
and bases, write  
equations for  
the following  
acid-base  
reactions and  
indicate each



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Basics

conjugate acid-

base pair: a)

$\text{HNO}_3 + \text{OH}^-$  (  
 $\text{H}_2\text{O} + \text{NO}_3^-$

$\text{HNO}_3$  and  $\text{NO}_3^-$

make one pair

$\text{OH}^-$  and  $\text{H}_2\text{O}$

make the other.

b)  $\text{CH}_3\text{NH}_2 + \text{H}_2\text{O}$

(  $\text{CH}_3\text{NH}_3^+ + \text{OH}^-$

---

Acid and Base

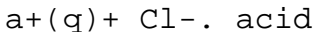
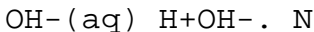
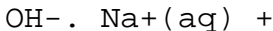
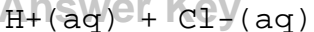
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acid salt acid

base acid.

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11.5b. Acids and bases can also be identified using an operational definition. Operational definitions are simply a list of properties.

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terms acid and  
base referred to  
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substances according to their acidic (sour) or basic (alkaline or bitter) properties dates back to ancient times. An acid was something with a sour taste, such as lemon juice, and a base was

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Something with a bitter taste, such as tonic water.

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Formula for each  
of the acids  
listed below: 1.  
Nitric acid  $\text{HNO}_3$   
2. Chloric  
acid  $\text{HClO}_3$   
3. Acetic acid  $\text{HC}_2\text{H}_3\text{O}_2$   
4. Hydrobromic acid  
 $\text{HBr}$   
5. Sulfurous  
acid  $\text{H}_2\text{SO}_3$   
6. Chlorous acid  
 $\text{HClO}_2$   
7. Hydrochloric

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acid HCl 8.

Phosphoric acid

$H_3PO_4$  9.

Nitrous acid HNO

$NO_2$  10.

Hydrofluoric

acid HF 11.

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$HClO_4$  12.

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A- Solutions 2,  
5 and 6 are  
bases, solution  
3 is an acid and  
solutions 1 and  
4 are salts. B-  
Solutions 2, 5  
and 6 are bases,  
solution 3 is an  
acid and

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Solutions 1 and 4 are distilled water. C- Solutions 2, 5 and 6 are bases, solution 3 is an acid, solution 1 is a salt and solution 4 can not be classified.

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acids include:

1. Giving foods a sour taste
- 2.

Forming  
solutions that  
conduct  
electricity ...

All the acids  
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following:

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hydroxide

Hydrobromic acid  
carbonic acid I

can define and  
identify

Arrhenius acids  
and bases. Label  
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base (B) or  
neutral (N).  
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